

## SYNTHESIS AND STUDIES ON THE CHEMISTRY OF CYCLOHEXAMINE - N-[(PHENYL) METHYLENE ] AND - N-[( FURFURYL) METHYLENE ] - N – OXIDES

S.Sakthi Raadha,

Professor, Department of chemistry, Adhiparasakthi Engineering College, Chennai, India  
sakthi09\_raadha@yahoo.com

### ABSTRACT

Synthesis of cyclohexamine - N – [(phenyl) methylene] and N – [(furfuryl) methylene] – N – oxide with NBS. On comparing both the  $^1\text{H}$  NMR and IR spectra of this brominated nitron with unbrominated nitron.

**Keywords:** *Infrared spectroscopy,  $^1\text{H}$  Nuclear magnetic resonance Spectroscopy, N-Bromosucinimide, Deuterated Chloroform.*

### INTRODUCTION

Considerable development in the knowledge and understanding of nitrones is mostly due to their increasing use in cycloaddition reactions, natural product synthesis, biologically active compounds and free radical traps. The existence of dipolar character in the nitron function has played a vital role in the development of newer synthesis approach to the formation of carbon-carbon bond. Nitrones are characterized by E-Z notation rather than the Syn-anti nomenclature. E-Z isomerization was observed to be fast at high temperature. It was found that aldonitrones exist exclusively in the Z- form unless a ring closure is encountered that forces the existence to be in the E- form. Nitrones are synthesised by condensation of carbonyl compounds with N – Nanosubstituted hydroxylamines. Alkylation of oximes at oxygen results in oxime ethers, whereas nitrogen results in nitrones . Recently few methods have been reported for the synthesis of nitrones by tungsten catalyst. Oxidation of secondary amines with hydrogen peroxide has been reported to give the corresponding nitrones in a single step and in excellent yields. The other recent report being the reaction of aryl or alkyl nitron compounds with 2 – butyl magnesium chloride to synthesis a new class of nitrones . Nitrones also obtained from oxaziridines and oxidation of schiff's bases.

Nitrones crystallite from anhydrous medium readily and also as adducts with various metal halides. Nitrones best dissolve in polar solvents. Very often they are hygroscopic and sufficiently basic to ferric with strong acids and adducts with Lewis acids. Alkylation at oxygen is readily achieved by heating with alkylating agents such as dimethylsulphate in an inert solvent or using a powerful alkylating agent such as triethyloxonium fluoroborate , to give corresponding o-alkylated salts. Reactions of nitrones with acylating agents such as acid chlorides, acid anhydrides,  $\text{PCl}_5$  ,  $\text{PCl}_3$  ,  $\text{PoCl}_3$ ., is more complex and tends to result in rearrangement to a secondary or tertiary nitrones.

## EXPERIMENTAL METHODS

Melting points are uncorrected. The IR Spectra of the samples were recorded in Perkin Elmer IR 577 instrument. <sup>1</sup>H NMR Spectra were recorded in CDCl<sub>3</sub> using 90 MHz R-32 Perkin Elmer instrument with TMS as the internal standard. UV Spectra were recorded in PU 8800 UV/ Vis Spectraometer (Philips) using ethanol as solvent.

### Synthesis of Cyclohexamine -N- [2/4 -Substituted Phenyl) Methylene] - N- oxide

(a) Benzaldehyde, and 2/4 -substituted benzaldehyde were purchased from 'ALDRICH 'and used as such.

#### (b) Preparation of Cyclohexamine oxime:

Hydroxylamine hydrochloride (2.5 g) is added and sodium acetate 4) in 10 ml H<sub>2</sub>O is taken in a small conical flask. The solution is warmed to about 40 ° and cyclohexane (2.5 g) is added. The vessel is stoppered securely with a cork and shaken vigorously for few minutes. The oxime separated out as a crystalline solid. This is cooled in ice. The crystals are filtered at the pump and washed with ice water. This is then crystallised from pet.ether (60-80 ° c) and dried. Yield: 2.6 g, m.p: 90 °

(c) Preparation of N - Cyclohexylhydroxylamine from cyclohexanone oxime using sodium cyanoborohydride/ acetic acid.

(d) To the stirring acetic acid (75 ml) at 25 ° c under nitrogen atmosphere are added cyclohexamine oxime (3.48 g, 30.9 mmole) and sodium cyanoborohydride (2.79 g, 45 mmole). The mixture is stirred for 3 h, and then water is added. The mixture is made basic with sodium hydroxide (pellets) and the resulting white solid is collected. The aqueous filtrate is extracted with ether and the ether extract is dried and concentrated in vacuo to afford a white solid. The combined solids are recrystallised from ether to give N - Cyclohexylhydroxylamine as colorless flakes. Yield: 2.88 g (81%), m.p: 138 - 139 ° c

### Synthesis of Cyclohexamine -N- [(Phenyl) Methylene] - N- oxide

Cyclohexylhydroxylamine ( 1.3 g , 0.01 mole ) obtained as above , benzaldehyde ( 1.28 g, 0.01 mole ) and benzene ( 60 ml ) are taken in a 250 ml R.B. flask. The mixture is heated for 1.5 h<sub>r</sub> on a steam bath and the resulting water is removed azeotropically. Evaporation of the remaining solvent and cooling of the residual oil resulted in Cyclohexamine -N- [(Phenyl) Methylene] - N- oxide as red prisms. This is recrystallised from pet.ether (40-60 ° c). Yield: 2 g (80.6 %), m.p: 81 - 83 ° c .

### Synthesis of Cyclohexamine -N- [ ( furfuryl ) Methylene ] - N- oxide

Cyclohexylhydroxylamine ( 1.560 g , 0.01 mole ) , furfural ( 0.96 g , 0.01 mole ) and alcohol (60 ml ) are taken in a 250 ml R.B flask. The mixture is refluxed for 1.5 h<sub>r</sub> on a steam bath. The solvent is then evaporated. On cooling Cyclohexamine -N- [(furfuryl) Methylene ] - N- oxide crystals are formed . It is recrystallised from pet.ether (40-60 ° C).

Yield: 1.4 g (73.2 %), m.p: 79 - 81 ° C

**RESULTS AND DISCUSSION****IR Spectra data:**

Frequency in $\text{cm}^{-1}$	Cyclohexamine-N- [(Phenyl)Methylene]-N-oxide	Cyclohexamine-N- [(furfuryl)Methylene]-N-oxide
C=N	1545	1580
N-O	1130	1140

 **$^1\text{H}$  NMR Spectral Data of Cyclohexamine - N-[(2/4 - Substituted Phenyl) Methylene] - N - Oxides:**  
( $\delta$  scale)

Postions	Cyclohexamine-N- [(Phenyl) Methylene]-N- oxide	Cyclohexamine-N- [(furfuryl)Methylene]-N- oxide
Alicyclicmethylene hydrogen	1.0-2.20, 10H, m	1.1-2.2 10H, m
C=N	3.95, 1H, m	3.85, 1H, m
$\alpha$ -H	7.70, 1H, s	7.75, s
$\text{C}_2$ and $\text{C}_6$ -H of aryl ring	8.55, 2H, d	-
$\text{C}_3$ , and $\text{C}_5$ -H of aryl ring	7.65, 3H, m	7.90, 1H, d, 7.60, 1H, d

**CONCLUSION:**

(i) These compounds show IR absorption  $1540 \pm 20\text{cm}^{-1}$  (due to C=N str.) and  $1135 \pm 15\text{cm}^{-1}$  (due to N-O str.) that are characteristic of nitron function. This together with the absence of carbonyl frequency around  $1751\text{cm}^{-1}$  confirms the formation of nitrones.

(ii) The  $^1\text{H}$  NMR spectra of all these compounds were characterized by their azomethine absorption signal 7.70 to 8.05 ppm apart from other signals due to the aromatic protons and other protons present in them. It can be noticed that when there is an electron withdrawing substitution in the para position of the  $\alpha$  - aryl ring, the azomethane is shifted to downfield region by approximately 0.2 ppm. In the case of electron releasing substitution at the para position, the same hydrogen is shifted to the up field region by 0.2 ppm.

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